

**activity (relative activity),  $a$** 

Defined by the equation

$$a = e^{\frac{\mu - \mu^0}{RT}}$$

where  $R$  is the gas constant,  $T$  the thermodynamic temperature,  $\mu$  the chemical potential and  $\mu^0$  the standard chemical potential the definition of which depends on the choice of the standard state.

**See also:** absolute activity

**Source:**

Green Book, 2nd ed., p. 49

PAC, 1996, 68, 957 (*Glossary of terms in quantities and units in Clinical Chemistry (IUPAC-IFCC Recommendations 1996)*) on page 989