Lewis acid

A molecular entity (and the corresponding chemical species) that is an electron-pair acceptor and therefore able to react with a Lewis base to form a Lewis adduct, by sharing the electron pair furnished by the Lewis base. For example:

\[ \text{Me}_3\text{B} + \text{NH} \rightarrow \text{Me}_3\text{B}^-\text{NH}_3 \]

Lewis acid  Lewis base  Lewis adduct

See also: coordination, dipolar bond

Source:
PAC, 1994, 66, 1077 (Glossary of terms used in physical organic chemistry (IUPAC Recommendations 1994)) on page 1135